BRIDGING CHEMISTRY

What is this booklet for:

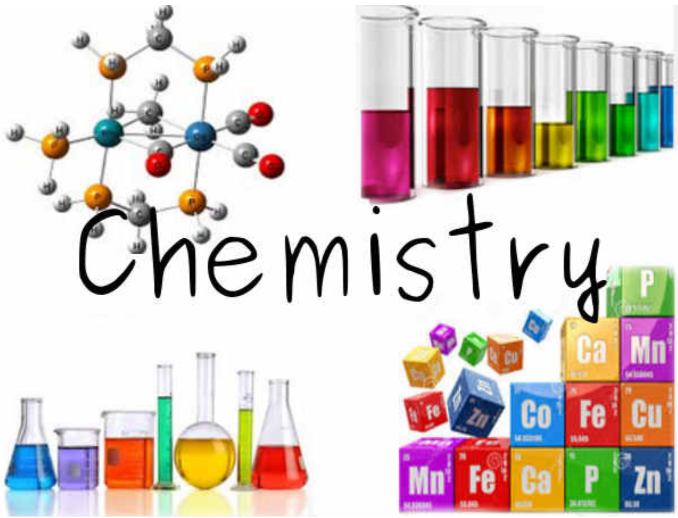
This is simply designed to be a *bridging* Chemistry booklet.

It has work to prepare you for the A level you are starting in September.

It contains a series of topics that you will have covered in GCSE and it is then extended into some A level standard work.

How to use the booklet:

1) Read over the explanation notes and examples
2) Look over work from your GCSE exercise books and revision guides
3) Look on the internet for other guidance, google the chapter titles!
4) COMPLETE the Tasks in the ANSWER booklet section.



General Guidance

You will need to know the basics as you soon as you start your AS chemistry lessons in September so make sure you arrive to your first lesson able to do the following tasks. This will allow you to focus on the skills required to master the more complex chemical concepts – giving you confidence rather than making you feeling like you are behind from the start! So come in to Yr 12 fresh but ready/fully prepared.

The way you study should change in Year 12, both in terms of the amount of independent study you do for each subject and the strategies you use/develop when studying – if it doesn't you are likely to be at risk of underperforming! You MUST keep on top of the workload from the start making regular summaries along the way and not leave revision until the end of year before the exams.

To show a positive attitude to learning in completing the following tasks you MUST....

- ✓ **avoid leaving gaps** - a big difference from GCSE to AS Level is how YOU take ownership for your learning. If you find a question difficult or challenging YOU must take action by researching the topic to help overcome any misunderstanding.

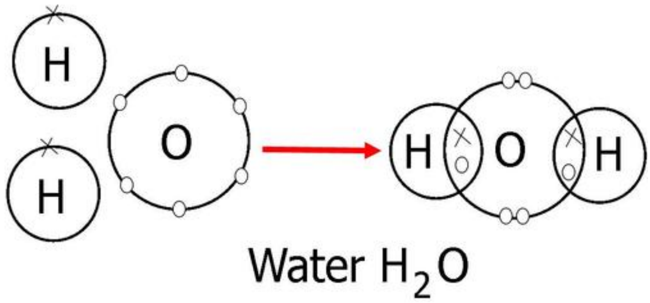
- ✓ **be thorough** – avoid cutting corners e.g. you MUST show full working in any calculations, never just give the final answer; write in full sentences so your work is meaningful during times of revision.

- ✓ **be independent** – there is a place for 'peer learning' but this can also limit your progress if you become too reliant on others to explain how to approach a question or regularly complete tasks working together. Make sure you try to overcome any barriers yourself first by being resourceful and carrying our further reading on difficult topics, then use your peers to check if you reached the same answer.

- ✓ **Prepare for lessons** – arrive to lessons ready to submit any work due in and refresh your memory of the work covered in the previous lesson by reading through your notes and possibly re-attempting one or two questions/tasks.

Example of a typical covalently bonded compound

Water



Chapter 1

Bonding

This is a cornerstone of chemistry, when elements react together they form new compounds which have two or more elements chemically joined.

There are two main types of chemical bond.

Ionic -----between a Metal and Non-metal

Covalent ------between Non-metal and Non- metal

Task 1

Decide if the compounds below are Ionically or covalently bonded together and why?

a) Ammonia NH_3

b) Zinc Oxide ZnO

c) Methane CH_4

d) Benzene C_6H_6

e) Potassium Dichromate $K_2Cr_2O_7$

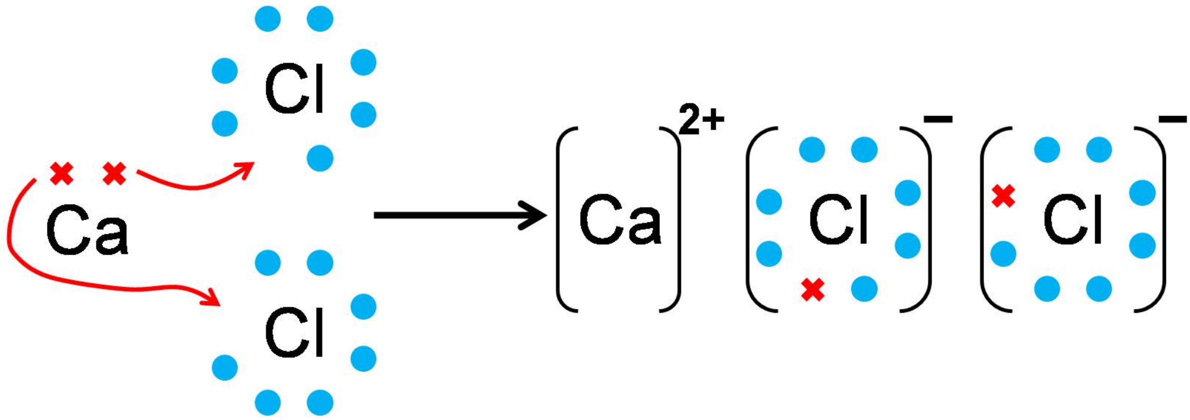
Ionic Bonding

This is an ELECTROSTATIC ATTRACTION between 2 oppositely charged species called IONS.

The compound is formed is neutral, which means it has no overall charge.

i.e. it has an equal amount of positive and negative charge from the different ions that are making it up.

How are IONS made?



This is seen by the diagram above:

METALS: (Calcium)	NON- METALS (Chlorine)
They form Positive ions as they lose their outer electrons to form a FULL OUTER SHELL.	They form NEGATIVE ions as they gain electrons to form a FULL OUTER SHELL.
Calcium 2 electrons in its outer shell as an element so LOSES 2 electrons to become a 2+ ion	Chlorine has 7 electrons in its outer shell so will GAIN 1 electron to become a 1- ion

Task 2

Draw out Atom and Ions for the following ionic compounds (like the calcium chloride diagram above)

1) Aluminium Oxide

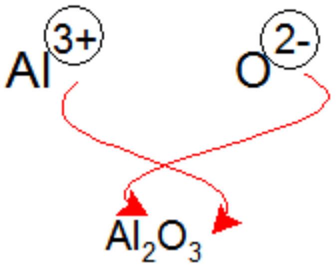
2) Lithium Oxide

3) Barium Nitride

Formula of Ionic compounds

When we form an Ionic compound we have oppositely charged ions attracted together so that a neutral compound is formed.

This means there is a balance between the positive metals ions and negative non-metal ions.

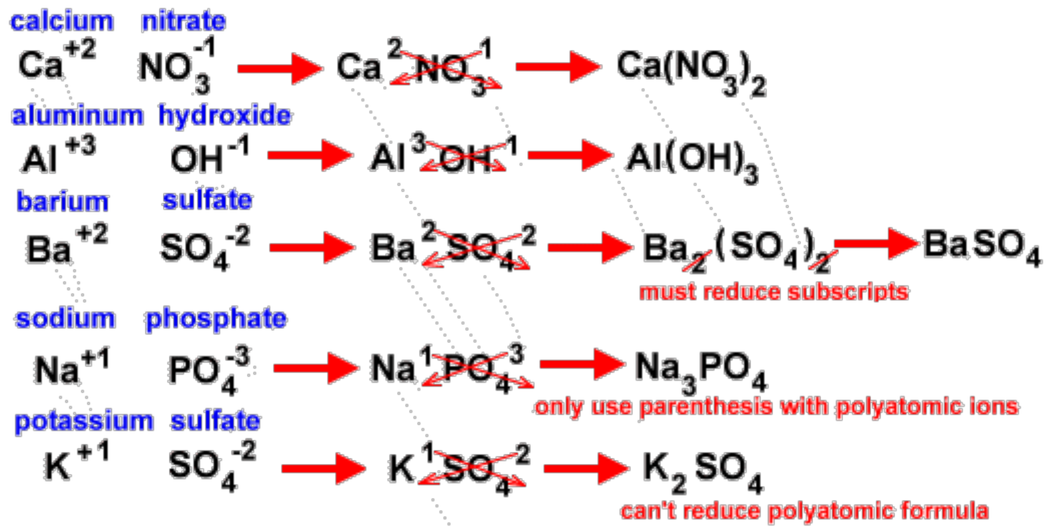


You swap the NUMBERS of the charge over

If a 1 you ignore it

If get 2 numbers the same ignore them

Aluminium Oxide made from Aluminium ions and Oxide ions.



Other examples above(don't worry about the writing in red)

Task 3 (Use appendix I)

Using the table of common ions work out the formula of the following ionic compounds.

1) Silver chloride
2) Lithium sulphate
3) Ammonium Hydroxide
4) Potassium Dichromate
5) Iron (II) Nitrate

Formula interpretation

When we have calculated the formula of a compound it is important we can interpret the information about the number of atoms and types of elements in the compound.

e.g.

Calcium Carbonate

$CaCO_3$

1 Ca

1 C

3 O

Task 4

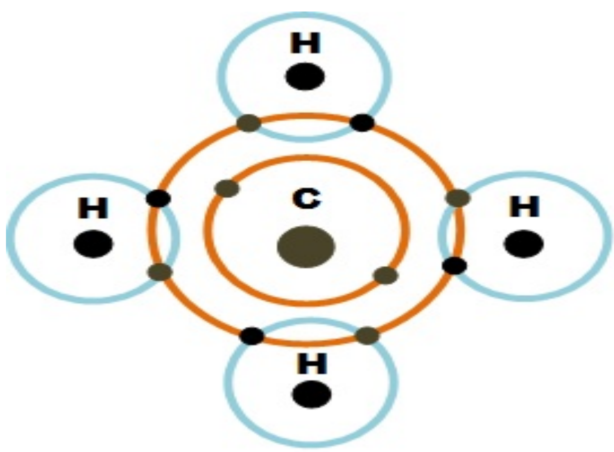
Look at the following compounds and work out the number and type of elements in the compound.

1) $AgNO_3$

2) $PbCO_3$

3) $SnCl_2$

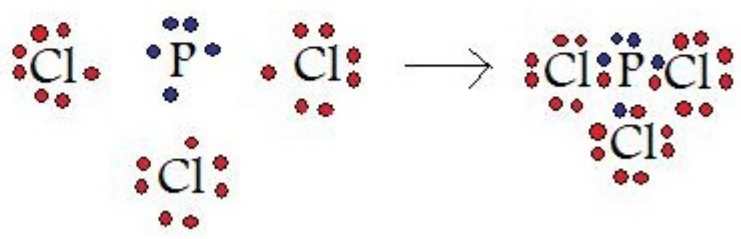
4) $Mg(OH)_2$



Covalent bonding

The covalent bond is made up from non-metal atoms that want to bond together.

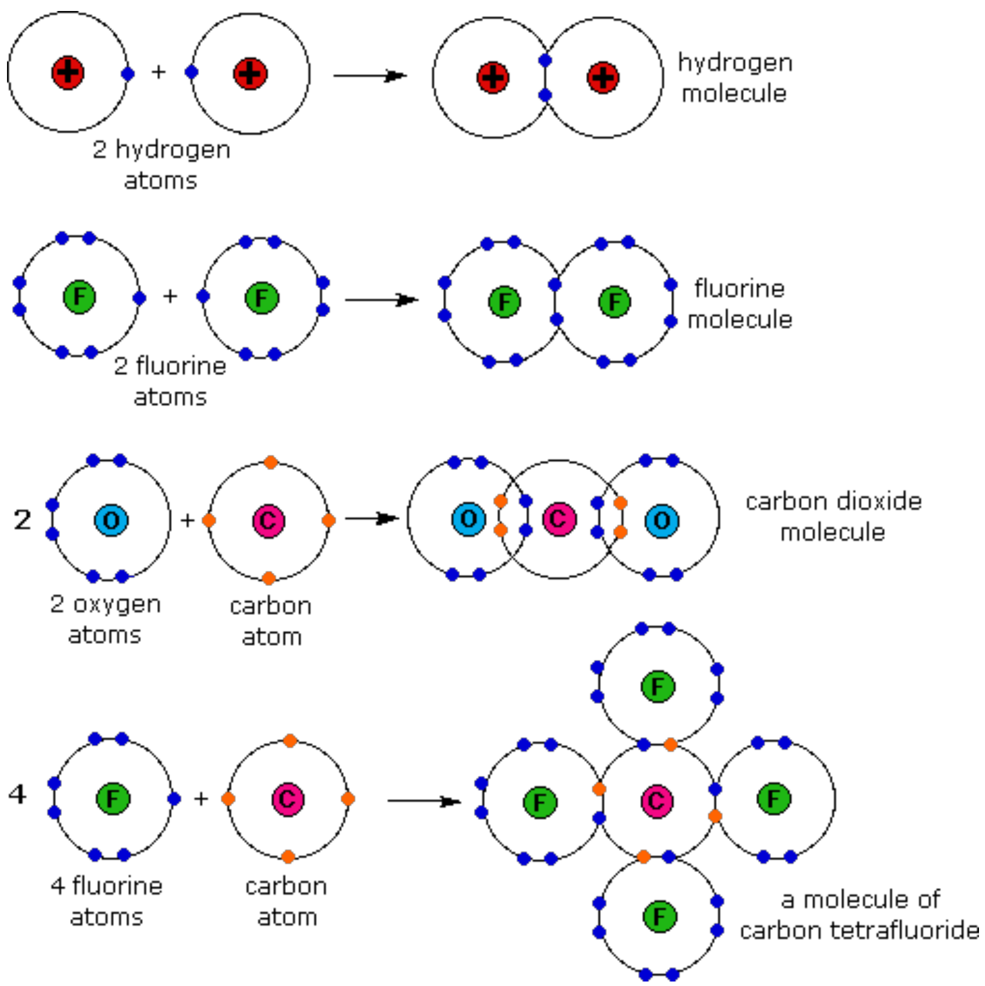
Covalent bonds are made from the atoms sharing their electrons to get a FULL OUTER SHELL.



The above example shows,

Phosphorus in group 5 with 5 outer electrons sharing 1 electron each with a chlorine atom which is in group 7.

Both the Phosphorus and Chlorine NOW have their FULL OUTER SHELL.



More examples

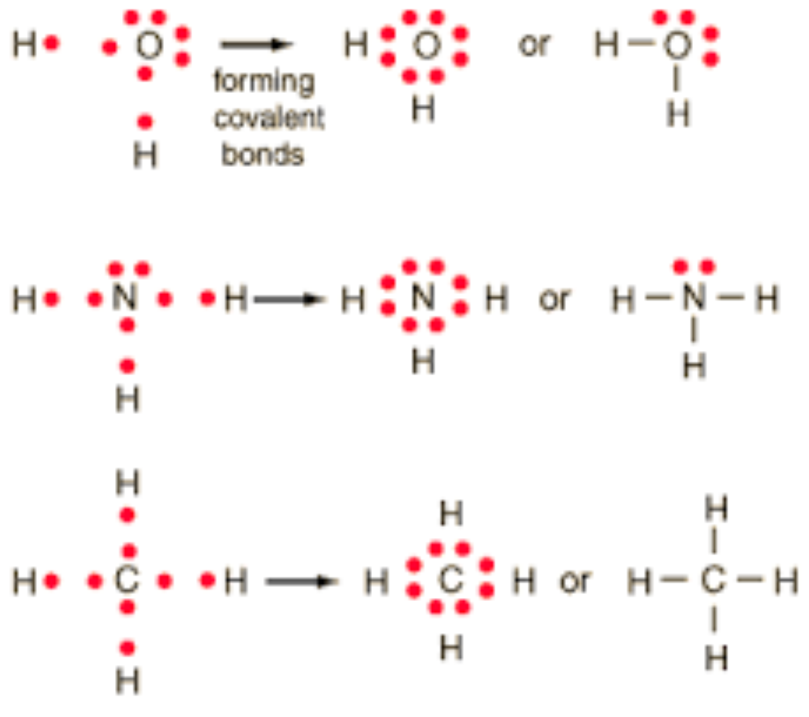
The example shows a series of covalently bonded molecules where the atoms have all got a FULL OUTER SHELL.
Please note
DOUBLE BOND on the CO_2 molecule .

The 4 SINGLE BONDS from the carbon attached to each individual F in the CF_4 molecule.

EXT Line diagrams

These are simpler versions of the shown DOT-CROSS diagrams where you show each bond (PAIR of ELECTRONS) as a line between the atoms in the molecule

e.g.



The extra pair of electrons that are not involved in the bonds are called LONE PAIR of electrons.

These are shown by the pair of 'dots' around the central atom.

Task 5
Draw out the Dot/ Cross diagrams and Line diagram of the following molecules:

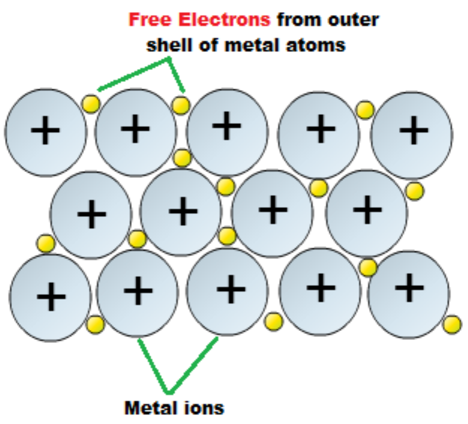
1) Ethane C_2H_6
2) Propene C_3H_6
3) Hydrogen Peroxide H_2O_2
4) Hydrogen Sulphide H_2S

Chapter 2
Structure

There are 4 main structures you need to be aware of
1) Metallic structure
2) Giant Ionic
3) Giant covalent / Macromolecular
4) Simple Molecular

1) Metallic

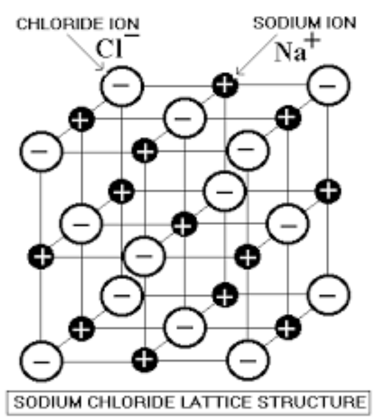
This occurs in metals.



These are strongly bonded structures which have HIGH boiling and melting points.
They CAN conduct electricity due to the FREE ELECTRONS.

2 Giant Ionic

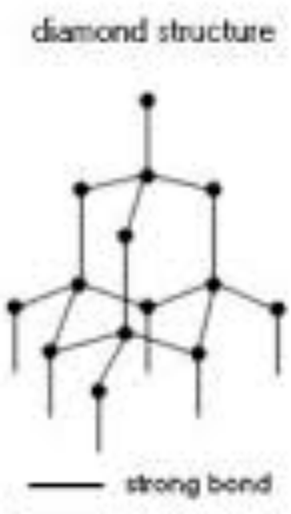
This occurs as a LATTICE of IONS electrostatically attached together with the positive ions being attracted to the negative ions.
It occurs in Ionically bonded compounds.

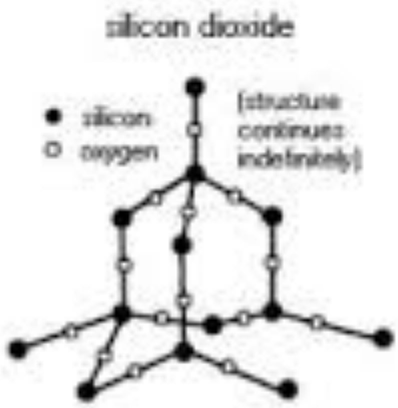


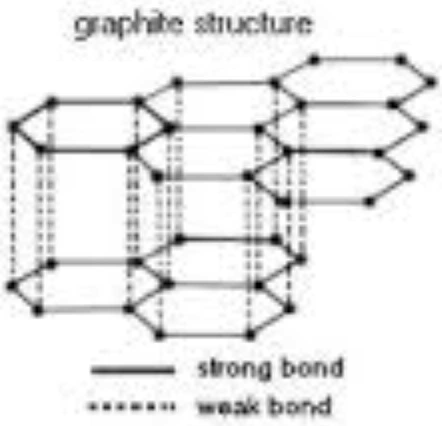
<u>3 Giant covalent / Macromolecular</u>

This occurs in a select number of covalently bonded compounds which have ALL their atoms covalently bonded together in a large structure.
Key examples are ALLOTROPES of carbon (look up what Allotrope means!) and silicon dioxide

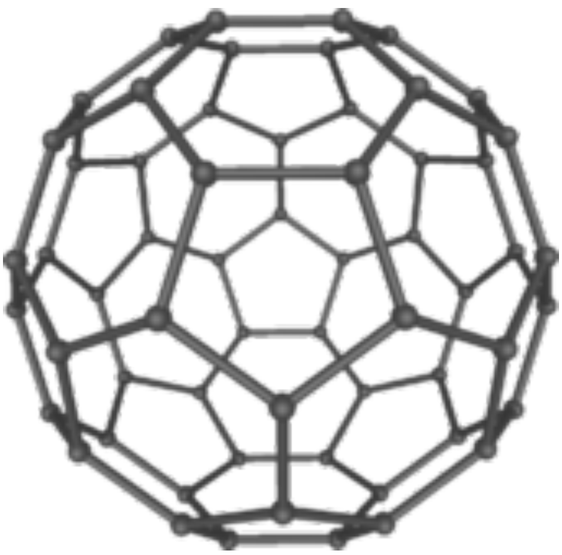
Diamond Silicon Dioxide Graphite







EXT
<u>Buckminster Fullerene</u>



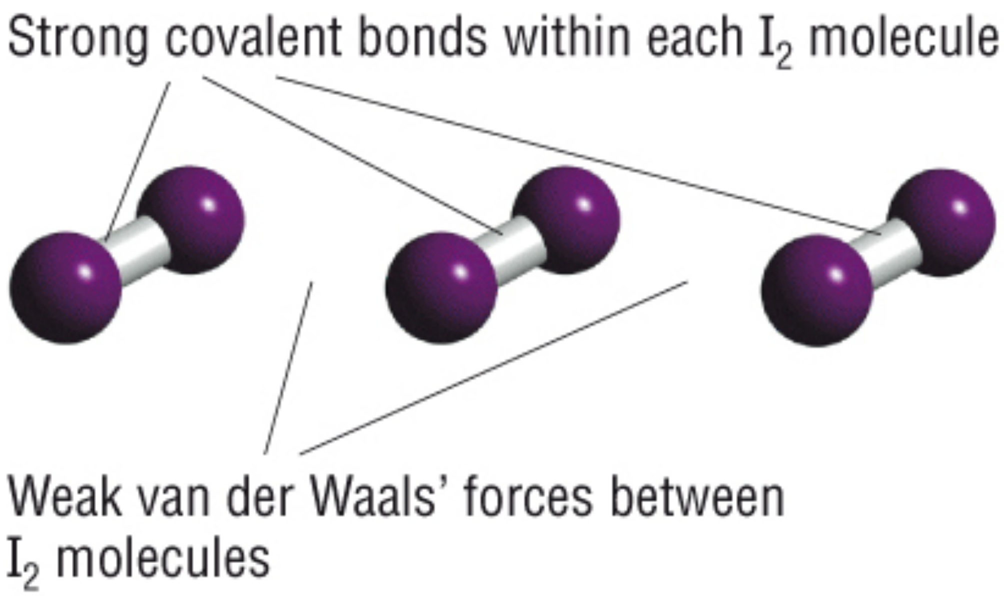
This is a C_{60} molecule in the shape of a football.

They were discovered in the UK in 1985 and the chemists involved won the Nobel prize in 1996.

4) Simple Molecular

This occurs in covalently bonded molecules which have STRONG covalent bonds inside the molecules
But
Much weaker INTERMOLECULAR bonds between the molecules.
The three types of INTERMOLECULAR bond/ force are:

- Van Der Waals
- Permanent Dipole
- Hydrogen Bond



Task 6
Research task
Find out what the trend in melting/ boiling point is for Na-Mg-Al (the metal in the third period)
Explain why there is this trend (linked to their structure)

http://www.creative-chemistry.org.uk/alevel/module1/trends8.htm

(basic source exemplar)

Chapter 3

Equations

We will be most interested in BALANCED symbol equations.
These show us exactly what elements are in the reactants and the products and we need the SAME amount on both sides of the equation.
Example

Calcium + Oxygen $\longrightarrow$ Calcium Oxide

$Ca + O_2 \longrightarrow CaO$

This is not balanced,
So we need to ADD large numbers in front of the formula given to balance it.

Firstly

$Ca + O_2 \longrightarrow 2\ CaO$

Added a 2 in front to get the right number of oxygen's.
But
This means we know have too many calcium's.

So we now need to add 2 on this side as well

$2Ca + O_2 \longrightarrow 2CaO$

It is now a Balanced equation.

Task 7

Balance the following equations:

1) $N_2 + H_2 \longrightarrow NH_3$
2) $CH_4 + O_2 \longrightarrow CO_2 + H_2O$
3) $Na + H_2SO_4 \longrightarrow Na_2SO_4 + H_2$
4) $SO_2 + NaOH \longrightarrow Na_2SO_3 + H_2O$
5) $C_2H_5OH + O_2 \longrightarrow CO_2 + H_2O$

State symbols

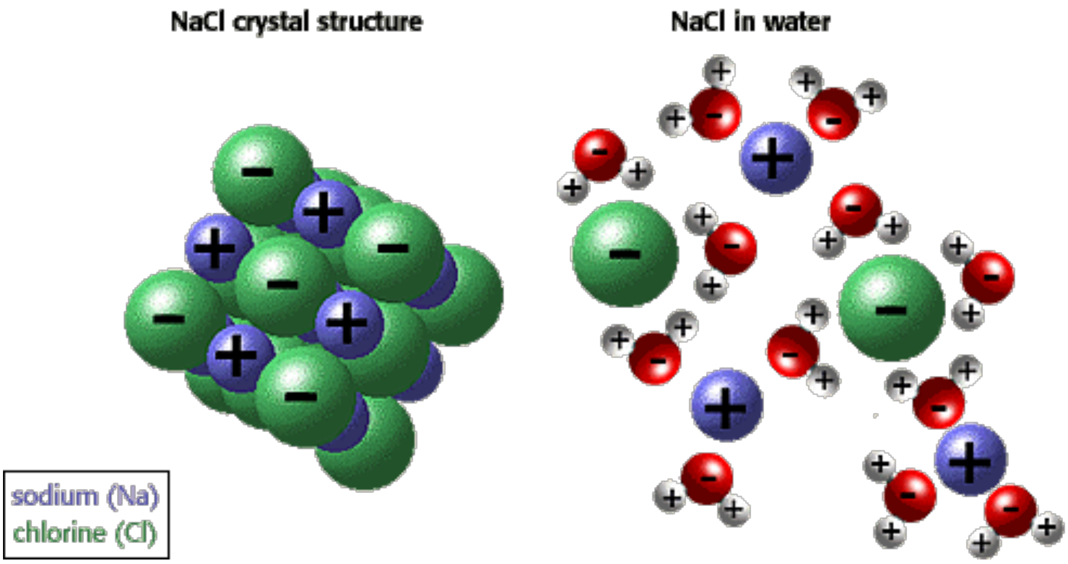
These are linked to the three states of matter

- Gas (g)
- Liquid (l)
- Solid (s)

Also we have (aq) for a solution.

EXT

Ionic compounds in solutions



When we dissolve an ionic compound it is the separate ions in the compound being split apart and bonded to the water.

$$NaCl\ (s) + aq \longrightarrow NaCl\ (aq)$$

Is in fact

Na^+ (aq)

Cl^- (aq)

Chapter 4

Mole work.

In its most basic form the 'MOLE' is just a word used to describe a number.

e.g.	Couple	2	
	Dozen	12	
	Mole	6.02×10^{23}	(602000000000000000000000)

Why this large number?

It was found that this number of ATOMS of any element is equal to the MASS NUMBER of this element in grams.

e.g.

6.02×10^{23} carbon atoms is equal to 12g

6.02×10^{23} neon atoms is equal to 20g

This leads to the FIRST mole equation.

$$\text{Moles} = \frac{\text{Mass}}{\text{R.A.M}} \quad \text{(relative atomic mass)}$$

e.g.

How many moles are there in 24g of carbon?

$$\text{Moles} = \frac{\text{Mass}}{\text{R.A.M}}$$

$$\text{Moles} = \frac{24}{12}$$

$$\text{Moles} = \text{2 moles of carbon}$$

Task 8

Calculate the number of moles in the following elements?

1) 59 g of cobalt
2) 4.14 g of lead
3) 1.08g of gold

This can get increased very quickly to include compounds and not just elements.

In this we use a very similar Mole equation:

$$\text{Moles} = \frac{\text{Mass}}{\text{R.F.M}}$$

This is the Relative formula mass

e.g. H_2O

H + H + O

1 + 1 + 16 = 18

e.g.
How many moles are their in 88g of carbon dioxide?

$$\text{Moles} = \frac{\text{Mass}}{\text{R.F.M}}$$

$$= \frac{88}{44}$$

$$= 2 \text{ mole}$$

CO_2

C + O + O

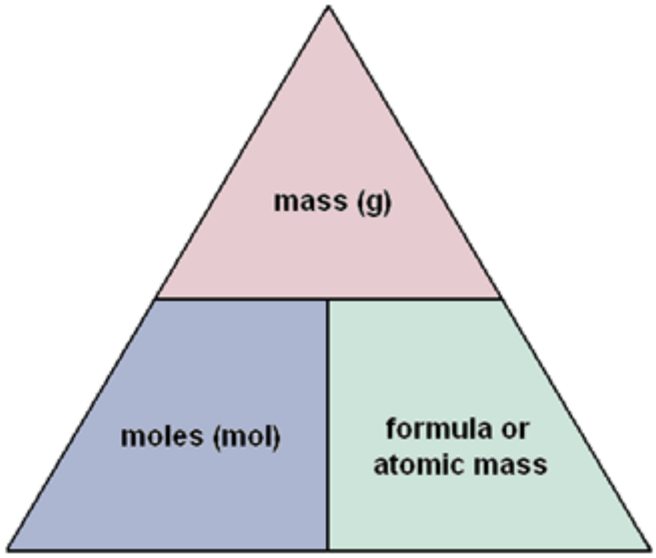
12 + 16 + 16 = 44

NOTE- Good practice
It is always good practice to start with the equation in word form then put the numbers in from the questions
It is also good practice to show how you have worked out the RFM so if there is an error you can still get method marks.

Task 9
How many moles are there in the following:
1) 62 g of sodium Oxide Na_2O
2) 174 g of lithium bromide LiBr
3) 3.2 g of oxygen
4) 1.24 g of Ammonia

Changing the equation

We can have this mole equation in a simple MAGIC TRIANGLE and easily change the aspect we are trying to work out.



So we may get asked to calculate the Mass or Relative formula mass.

Task 10

Calculate the :

1) Mass of 2 moles of calcium metal
2) 0.25 moles of lead carbonate $PbCO_3$
3) The formula mass of a compound which has 0.5 moles of mass 14g

EXT

Harder question

Task 11

250g of hydrated copper sulphate ($CuSO_4 .x\ H_2O$) is collected and a student want to calculate the number of moles of water attached to the copper sulphate, the x value.

The student completely dried the copper sulphate and the new mass was found to be 160g

a) Calculate the moles of copper sulphate
b) Calculate the mass of lost water
c) Calculate the number of moles of lost water
d) Therefore deduce the formula of the hydrated copper sulphate.

Moles and solution

When we dissolve a solid in water we create a solution.
We use a different mole equation to calculate the moles in the solutions we create.

Molarity / M
Mol/dm^3 ml or cm^3

$$\text{Moles} = \frac{\text{Conc} \times \text{Vol}}{1000}$$

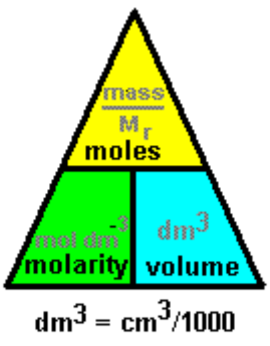
e.g.
How many moles are there in 250 cm 3 of 0.1 M Hydrochloric acid ?

$$\text{Moles} = \frac{\text{Conc} \times \text{Vol}}{1000}$$

$$= \frac{0.1 \times 250}{1000}$$

$$= \underline{0.025 \text{ Moles}}$$

This equation can again be moved around if you have to calculate the concentration using the moles and volume.



Task 12

1) Calculate the moles in 40 ml of 5M of sodium hydroxide solution
2) What is the concentration when you dissolve 2 moles of acid in 100ml of water
3) How many moles are there in 500ml of 0.1 mol/dm^3 of salt solution
4) What is the concentration of 0.25 moles of alkali in 25 ml

EXT
Combining our work

We often need to combine this work on moles and work out the mass of a solid we need to make up a set concentration of a solution.

I.e. we want to make 100ml volume of a 0.5 M solution of sodium Hydroxide, how much mass do we need to dissolve?

1) How many moles are in this solution,

$$\text{Moles} = \frac{\text{Conc} \times \text{Vol}}{1000}$$

$$= \frac{0.5\ M \times 100ml}{1000}$$

$= \underline{0.05}$ Moles of sodium hydroxide in solution

2) What mass do we need for that many moles,

NaOH
Na + O + H
23 + 16 + 1 = 40

Mass = moles x RFM

= 0.05 x 40

= <u>2 g</u>

So we will need to dissolve 2 g in the 100ml to make the required solution concentration of 0.5M.

Task 13

1) How many grams of potassium oxide (K_2O) are needed to make 100ml of a 0.5M solution ?
2) What is the concentration of a solution when we dissolve 730g of hydrochloric acid in 350 cm^3?
3) What is the mass of calcium oxide, CaO needed to make a 250 ml volume of 0.5 M solution?

NOTE- HINT

Keep looking carefully at the units

Ml= cm^3 for volume   mol/dm^3 = Molarity =M for concentration

<u>**Molar Ratio**</u>

This is the link between the balanced symbol equations and the amount of moles of each substance in the reaction.
Simply it is the ratio of the numbers in front of the compounds in the balanced symbol equation.

e.g.

$$2Ca \quad + \quad O_2 \longrightarrow 2CaO$$

In this equation the Molar ratio is:

$$2 \quad : \quad 1 \longrightarrow 2$$

Means:

2 moles of calcium will react with 1 mole of oxygen and we will make 2 moles of the calcium oxide.

As it is a ratio these numbers can be varied,

So if we actually had **10** moles of the calcium?

$$2 \quad : \quad 1 \longrightarrow 2 \qquad \text{original ratio}$$

10

$$10 \quad : \quad 5 \longrightarrow 10$$

So 10 moles of the calcium would react with 5 moles of the oxygen and form 10 moles of the calcium oxide

Or if we wanted to make **0.25** moles of the calcium oxide

$$2 \quad : \quad 1 \longrightarrow 2 \qquad \text{original ratio}$$

0.25

$$0.25 \quad : \quad 0.125 \longrightarrow 0.25$$

We would need 0.25 moles of the CaO

Final mole equation work

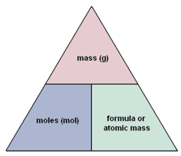
We are often asked to calculate how much we will produce in a reaction from a certain starting amount of reactants, or how much reactants we will need to make a set amount of products.

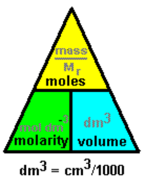
We put together the :

- Molar ratio work with the balanced equation
- The different moles equations

NOTE

If it involves a SOLID it is …





If it involves a solution it is ..

e.g.

Calcium oxide reacts with water to form calcium hydroxide.

$CaO + H_2O \rightarrow Ca(OH)_2$

If I started with 28g of the calcium oxide what mass of calcium hydroxide would I make, and if it was in 100ml of water what would its concentration be

1	: 1		1	Molar Ratio
CaO	+ H_2O	$\rightarrow$	$Ca(OH)_2$	
28g				

Moles = $\frac{\text{Mass}}{\text{RFM}}$

$= \frac{28}{56}$

=0.5 moles

$Ca(OH)_2$

Ca + O + H + O + H

40 + 16 + 1 + 16 + 1 = 74

0.5 0.5 0.5 *New molar ratio*

Mass = Moles x RFM

= 0.5 x 74

= 37g

And the solution concentration would be:

0.5 moles
100ml

$$\text{Conc} = \frac{1000 \times \text{mole}}{\text{Vol}}$$

$$\text{Conc} = \frac{1000 \times 0.5}{100}$$

Conc = 5 mol/dm^3

Task 14

1) Calcium cyanamide $CaCN_2$ reacts with water to form calcium carbonate and ammonia

   $CaCN_2 + 3H_2O \rightarrow CaCO_3 + NH_3$

   What mass of calcium carbonate is formed if 20g of the $CaCN_2$ is reacted with excess water.

2) Magnesium burns in air to make magnesium oxide

   $2Mg + O_2 \rightarrow 2MgO$

   What mass of magnesium would you need to create 0.8g of magnesium oxide powder.

3) Iron reacts with water to form iron oxide and hydrogen

   $3Fe + 4H_2O \rightarrow Fe_3O_4 + 4H_2$

   If the student starts with 1.68g of iron and it undergoes a complete reaction

   i) Number of moles of iron started with?
   ii) Moles of tri Iron oxide formed
   iii) Mass of tri iron oxide formed
   iv) The concentration of this solution if we had 500ml of water in the reaction?

4) 25 ml of 0.1 M HCl reacts with 50ml of NaOH solution fully

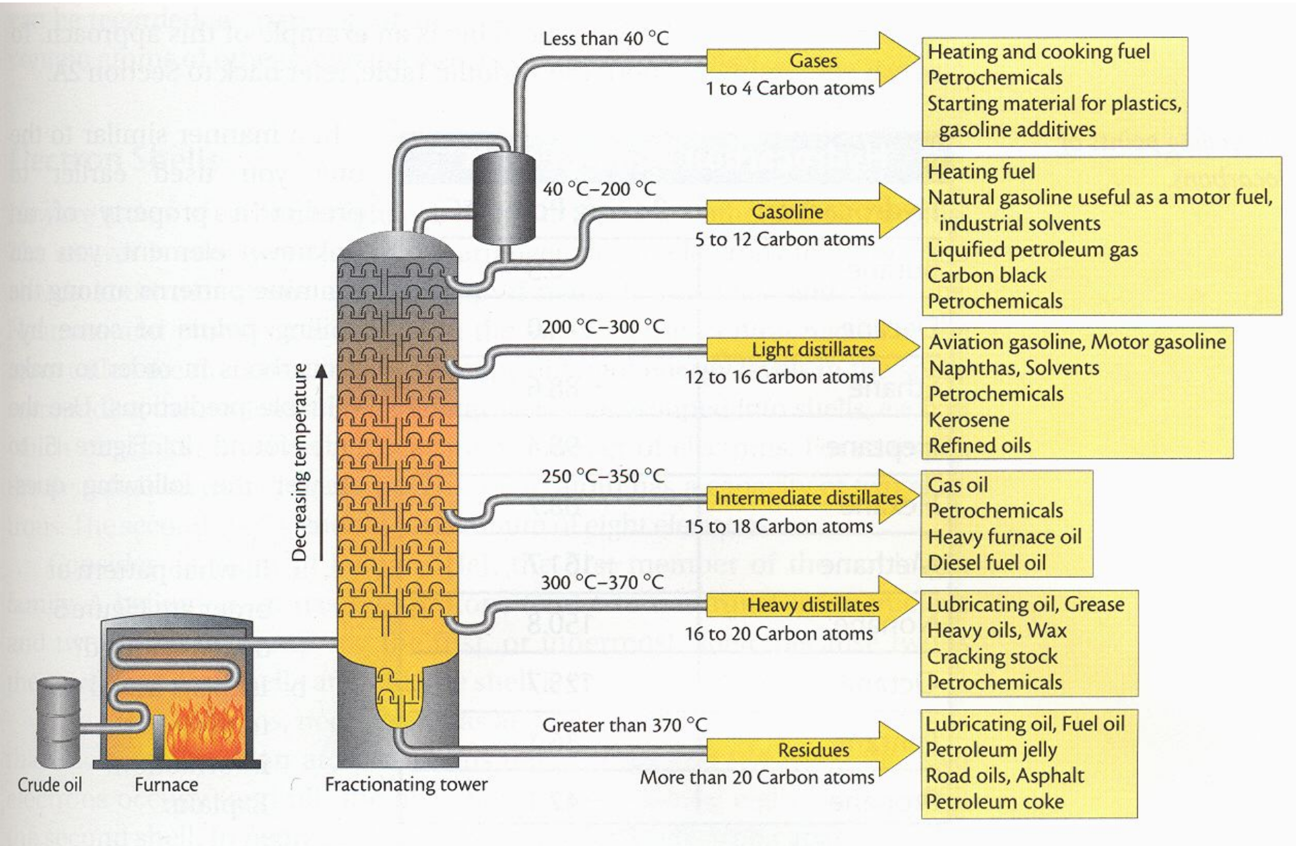
   What is the concentration of the NaOH solution.

   $HCl + NaOH \rightarrow NaCl + H_2O$

Chapter 5

Organic chemistry

This is a branch of chemistry that looks at compounds of carbon chained molecules.
The main source of these compounds is CRUDE OIL.
We FRACTIONALLY DISTILL this to separate it out into different FRACTIONS which have similar boiling points, size and properties.



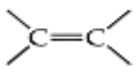
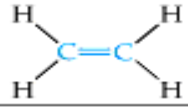
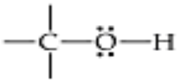
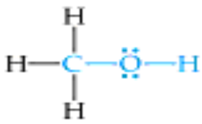
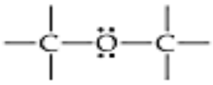
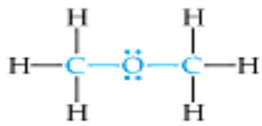
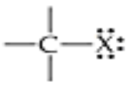
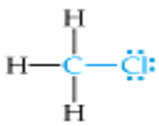
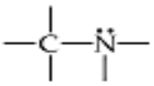
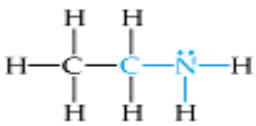
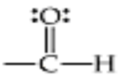
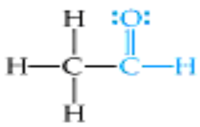
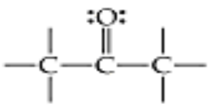
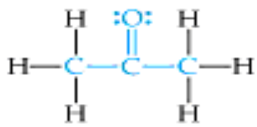
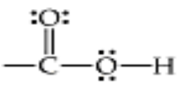
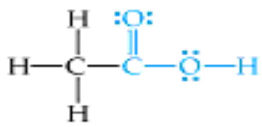
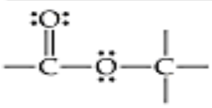
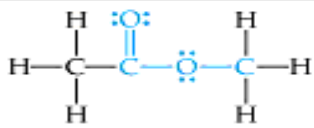
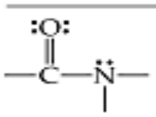
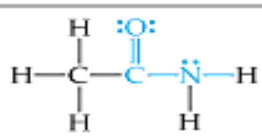
Task 15

Imagine you are a small CH_4 molecule in crude oil and you are being fractionally distilled,
What happened to you?
Why?
What happens to other molecules at the same time?
Why?
USE correct technical language to explain what's going on.

Types of organic compound

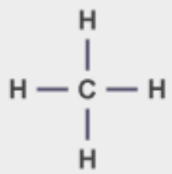
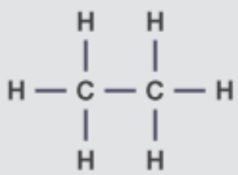
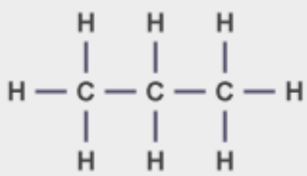
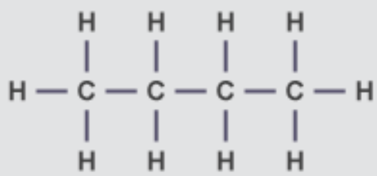
There are lots of different types of organic compound which are based upon their FUNCTIONAL GROUPS or parts of the compound which determine how they react.

TABLE 25.4 Common Functional Groups in Organic Compounds

Functional Group	Type of Compound	Suffix or Prefix	Example	Systematic Name (common name)
	Alkene	*ene*		Ethene (Ethylene)
	Alkyne	*-yne*		Ethyne (Acetylene)
	Alcohol	*-ol*		Methanol (Methyl alcohol)
	Ether	*ether*		Dimethyl ether
 (X = halogen)	Haloalkane	*halo-*		Chloromethane (Methyl chloride)
	Amine	*-amine*		Ethylamine
	Aldehyde	*-al*		Ethanal (Acetaldehyde)
	Ketone	*-one*		Propanone (Acetone)
	Carboxylic acid	*-oic acid*		Ethanoic acid (Acetic acid)
	Ester	*-oate*		Methyl ethanoate (Methyl acetate)
	Amide	*-amide*		Ethanamide (Acetamide)

The table shows the most common functional groups with examples and naming ideas.

Another aspect of organic compounds is the SERIES (called HOMOLOGOUS SERIES) you have of compounds which all have the same functional group. These all increase by –CH2- each time and have a common pattern of naming linked to the number of carbons in the compound.

Name	Molecular formula	Full structural formula
Methane	CH_4	
Ethane	C_2H_6	
Propane	C_3H_8	
Butane	C_4H_{10}	

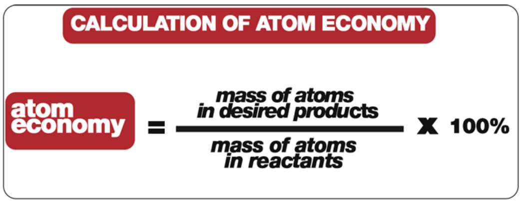
<u>Task 16</u>
Research
What are the FIRST 10 stem names for organic compounds
using alcohols as an example write out the molecular formula for the first 10, draw out the full structural/ displayed formula for the first 10 and names then as well.
(HINT complete a table like one above but for the first 10 alcohols!)

Chapter 6

Calculations on efficiency of reactions.

There are two main methods that are used to look over the efficiency of chemical reactions.

1) Atom economy



This is a measure of the useful products compared to all the products.
e.g.
Ethanol is decomposed into useful ethane and waste water.

	Ethanol	$\rightarrow$	Ethene	+	Water
	C_2H_5OH	$\rightarrow$	C_2H_4	+	H_2O
RFM	46		28		18

$$\text{Atom economy} = \frac{\text{mass of useful product}}{\text{mass of all reactants}} \times 100$$

$$= \frac{28}{46} \times 100$$

$$= \underline{\mathbf{60.9\%}}$$

Task 17

What is the Atom economy in:

1) Hydrogen is used in synthesising ammonia and is made on a large scale from reacting methane with water

methane + water ==> hydrogen + carbon monoxide

$$\mathbf{CH_4 + H_2O \implies 3H_2 + CO}$$

2) In the blast furnace where we form Iron .

$$\mathbf{Fe_2O_{3(s)} + 3CO_{(g)} \Longrightarrow 2Fe_{(l)} + 3CO_{2(g)}}$$

2) Percentage yield

This is the second method we use to calculate the efficiency of the reaction. This gives an idea of what is actually formed in reality as compared to what we would expect to be formed.

$$\text{Percent Yield} = \frac{\text{Actual Yield}}{\text{Theoretical Yield}} \times 100\%$$

NOTE

You are often given the actual amount you form BUT you have to work out the theoretical amount from a mole calculation.

e.g. Ethanol is decomposed into useful ethane and waste water.

Ethanol $\longrightarrow$ Ethene + Water

$C_2H_5OH \longrightarrow C_2H_4 + H_2O$

We create 1.4 g of the ethene from a starting mass of 4.6g of ethanol, what is the percentage yield.

CALC $\text{Moles} = \frac{\text{Mass}}{\text{RFM}}$

$\text{Moles} = \frac{4.6}{46}$

= 0.1 moles   0.1 moles : 0.1 moles

Mass = Moles x RFM

= 0.1 x 28

= 2.8 g

This is the theoretical yield amount

i.e this is the full amount that could possibly be formed

Final calc $\text{percentage yield} = \frac{\text{Actual}}{\text{Theoretical}} \times 100$

$= \frac{1.4}{2.8} \times 100$

= 50%

Task 18

1) When 5.00 g of $KClO_3$ is heated it decomposes according to the equation:

$2KClO_3 \rightarrow 2KCl + 3O_2$

a) Calculate the theoretical yield of oxygen.

b) Give the % yield if 1.78 g of O_2 is produced.

c) How much O_2 would be produced if the percentage yield was 78.5%?

2) The electrolysis of water forms H_2 and O_2.

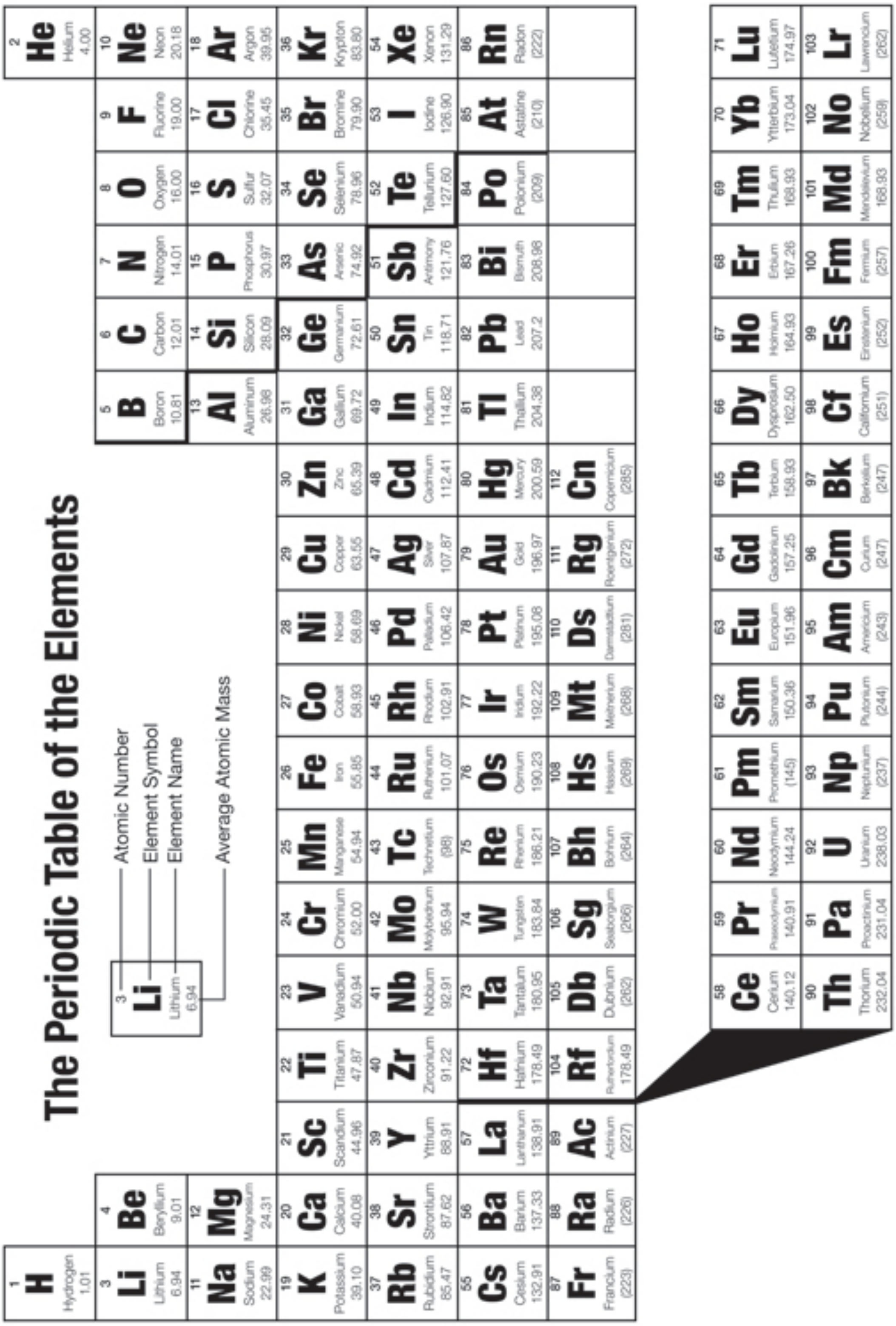
$2H_2O \rightarrow 2H_2 + O_2$

What is the % yield of O_2 if 12.3 g of O_2 is produced from the decomposition of 14.0 g H_2O?

Appendix I
Common ions

Positive Ions (cations)		Negative Ions (anions)	
Name	Formula	Name	Formula
Hydrogen	H^{+}	Chloride	Cl^{-}
Sodium	Na^{+}	Bromide	Br^{-}
Silver	Ag^{+}	Fluoride	F^{-}
Potasssium	K^{+}	Iodide	I^{-}
Lithium	Li^{+}	Hydroxide	OH^{-}
Ammonium	NH_4^{+}	Nitrate	NO_3^{-}
Barium	Ba^{2+}	Oxide	O^{2-}
Calcium	Ca^{2+}	Sulphide	S^{2-}
Copper(II)	Cu^{2+}	Sulphate	SO_4^{2-}
Magnesium	Mg^{2+}	Carbonate	CO_3^{2-}
Zinc	Zn^{2+}	Hydrogencarbonate	HCO_3^{-}
Lead	Pb^{2+}		
Iron(II)	Fe^{2+}		
Iron(III)	Fe^{3+}		
Aluminium	Al^{3+}		

Appendix II



Periods are Horizontal across the table

Groups are vertical down the table

Group1 –Alkali metals   7 Halogens   8 Nobel Gases